

Dedicated to the birth centenary of the Father of Nation, Bangabandhu Sheikh Mujibur Rahman

Invited Review Article

MICROSCOPIC ORIGIN OF IMMISCIBILITY AND SEGREGATION IN LIQUID METALLIC BINARY ALLOYS

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ABSTRACT

Microscopic description in the study of immiscibility and segregating properties of liquid metallic binary alloys has gained a renewed scientific and technological interests during the last eight years for the physicists, metallurgists and chemists. Especially, in understanding the basic mechanisms, from the point of interionic interaction, and how and why segregation in some metallic alloys takes place at and under certain thermodynamic state specified by temperature and pressure. An overview of the theoretical and experimental works done by different authors or groups in the area of segregation combining electronic theory of metals, statistical mechanics and the perturbative approach is presented in this review. Main attention in this review is focused on the static effects such as the effects of energy of mixing, enthalpy of mixing, entropy of mixing and understanding the critical behavior of segregation of alloys from the microscopic theoretical approach. Investigation of segregating properties from the dynamic effects such as from the effects of shear viscosity and diffusion coefficient is just becoming available. However, we have restricted this review only on static effects and their variation of impacts on different alloys.

Keywords: *Segregation, Thermodynamics of mixing, electronic theory of metals, Critical temperature and critical concentration, Perturbative approach.*

1. INTRODUCTION

Some advancement in understanding the segregating properties, miscibility gap, demixing tendency etc. of some metallic binary alloys, has been made, so far, from the empirical models [1-3], and phenomenological theories [4,5] in conjunction with arbitrary concepts of association and dissociation [6-8]. Experimental data for some liquid segregating alloys [9-20] play the pivotal role to arouse the interest in theoretical study, in particular to understand the critical behaviors. This knowledge is required to find the possible application of segregating materials to innovate technology and to industries for car engines, electrical contacts and switches, separation of impurities from the iron melts, ceramic industries, cosmetic and the food industry.

Known signatures of the existence of immiscibility, segregation, miscibility gap, and critical properties of segregating alloys are deviation from Rault's law, concave downward of the free energy of mixing profile [21-25], concave upward of entropy [24,25] and enthalpy of mixing [25], large density fluctuation displayed by concentration-concentration structure factors [26], large difference in partial coordination numbers [27] derived by using partial pair correlation functions $g_{ii}(r)$ and g_{12} , sudden sharp bending of the atomic transport properties as a function of concentration [25], positivity of short range order parameter [28- 32], exhibiting some sort of scaling laws [25] etc. But, understanding of the actual mechanisms involved behind these signatures is a great challenge to physicists, metallurgists and the material engineers.

Segregating properties of liquid binary alloys may be studied microscopically from the static [21-24] and dynamic effects [25]. The static effects may be observed from the thermodynamic properties of mixing, coordination number derived from structural properties etc. The dynamic effects can be seen from the atomic transport properties such as coefficient of viscosity and diffusion coefficient, and the electronic transport properties such as electrical resistivity [26]. In this review, a microscopic theoretical approach that involves the electronic theory of metals [23,33-35], perturbation theory [36,38], the hard sphere reference system [39,40] and the statistical mechanics is employed. Electron ion interaction is described by a local pseudopotential, the interionic pair interaction is derived from the energy band structure which is finally employed to evaluate static structure of liquid metals and their alloys. Specifically, the form factors of the pseudopotential is used to find the effective pair potentials and the volume dependent contribution to the free energy. The knowledge of pair potentials is essential to have pair correlation functions, the energy of the reference system and that of the attractive tail [33,37].

There are many liquid metallic binary alloys which exhibit miscibility gap or segregation at certain thermodynamic state. Some of these alloys are Li-Na, Al-In, Al-Pb, Al-Bi, Zn-Bi, Bi-Ga, Ga-Pb, Ga-Hg, Pb-Sn, Fe-Cu, Co-Cu, Cu-Pb etc. Of them, only a few systems such as Al-In, Al-Bi, Zn-Bi, Fe-Cu, Co-Cu are systematically studied employing microscopic theory, of course, empirical or phenomenological theories are applied to study some other systems [26]. For Al-In, Al-Bi, Zn-Bi, Fe-Cu, Co-Cu liquid binary alloys break down details are available (see below), from which one can analyze which component of the interionic interaction contributes how much or dominates in making the segregation to happen. Finally, comparison of these results with the experimental data would help understand the origin of segregation from the microscopic point of view and also the limitations of the employed theoretical approaches. Very little efforts have been spent so far in the study of immiscibility or segregating behavior of liquid metallic alloys from the effects of dynamic properties when it is compared with that of the static effects. So, it demands further to have considerably more studies in this direction. For this purpose, the easiest way is to invoke the Rice-Allnatt theory for atomic transport properties. Because, analytic expressions for shear viscosity and diffusion coefficient are already available for elemental [41,42] and binary alloys [43].

The layout of this review is as follows. Relevant theories are briefly discussed in section 2. Section 3 is devoted to the results obtained from the empirical and phenomenological theories. Results for the segregating properties for different alloys calculated from the microscopic theories are presented in section 4. A brief comparison of the impacts of the interionic interactions on different alloy systems is done and analyzed in the concluding section 5.

2. THEORY

Different theories relevant to the present review article are briefly presented below.

2.1 *Thermodynamic relations involved*

A macrostate of a condensed system may be described by four independent variables. These are pressure, p , volume, V , temperature, T , and entropy, S . Here, p and V form a pair representing the mechanical degrees of freedom, and T and S form another pair representing the thermal degrees of freedom. Any two of the four variables may be chosen in six different ways. Of them four pair of variables are (p,T) , (p,s) , (V,T) and (V,S) , each of which contains one variable from the mechanical and another from the thermal degrees of freedom. Thermodynamic functions constructed by these

pairs are Gibbs free energy, $G(p,T)$, enthalpy, $H(p,S)$, Helmholtz free energy, $F(V,T)$, and the internal energy, $E(V,S)$.

The free energy in thermodynamics is the amount of energy of the system free to work. Internal energy of a system is the sum of kinetic energy, potential energy, rotational energy and the vibrational energy etc. Of course, in the magnetic systems the magnetic energy [44,45] and for a finite sized sample the surface energy correction to be counted in the above functions [46]. For a monoatomic system (also in random binary alloys) there are no rotational and vibrational energy contribution in general. However, the above thermodynamic functions are not independent to each other. They are rather interconnected. The Helmholtz free energy (for the bulk) is (dropping the variables for brevity)

$$F = E - TS. \quad (1)$$

The Gibbs free energy

$$G = F + pV = E - TS + pV. \quad (2)$$

The enthalpy

$$H = E + pV. \quad (3)$$

And the change in the internal energy

$$dE = TdS - pdV. \quad (4)$$

From the theoretical point of view, we can further analyze the above relations, for example, for zero pressure, i.e., at $p = 0$,

$$G = F \quad (5)$$

$$H = E. \quad (6)$$

Again, most of the experimental data for thermodynamic quantities available in the literature are at standard temperature and pressure. In one

atmospheric pressure the value of the product of pV appears to be very small when compared with other terms of the above thermodynamic functions. So, in one or two atmospheric pressure or less one can write

$$H \approx E ; G \approx F. \quad (7)$$

We note here that other physical quantities such as heat capacity, compressibility etc. can be derived from equations (1) to (3) [47-49].

For binary alloys the free energy of mixing is

$$\Delta F = F_{\text{alloy}} - \sum_a C_a F_a \quad (8)$$

where F_{alloy} is the free energy of the alloy, C_a is the concentration of the a -th component and, F_a is the free energy of the a -th element in the same thermodynamic state. Similarly, the enthalpy of mixing may be expressed as

$$\Delta H = H_{\text{alloy}} - \sum_a C_a H_a \quad (9)$$

and the entropy of mixing as

$$\Delta S = S_{\text{alloy}} - \sum_a C_a S_a \quad (10)$$

It is worth noting that, the thermodynamics is a phenomenological subject because all relations in thermodynamics are obtained just looking at the experimental results. The only way to have microscopic description of the thermodynamic quantities is through the statistical mechanics [48,49].

3. MICROSCOPIC THEORY FOR METALLIC SYSTEMS

3.1 The pair correlation function

Let us consider N ions each of valence Z are there in a volume V in a liquid metallic system. So, the total number of conduction electrons in this system is NZ . The Hamiltonian of the sample may be written as

$$\begin{aligned} H &= H_e + H_{ee} + H_{ei} + H_i + H_{ii} \\ &= \sum_{i=1}^{NZ} \frac{p_i^2}{2M} + \frac{e^2}{2} \sum_{i \neq j}^{NZ} \frac{1}{|\vec{R}_i - \vec{R}_j|} + \sum_{i,l} v(|\vec{R}_i - \vec{r}_l|) \\ &\quad + \sum_{l=1}^N \frac{P_l^2}{2m} + \frac{1}{2} \sum_{l \neq l'}^N w(|\vec{r}_l - \vec{r}_{l'}|) \end{aligned} \quad (11)$$

where H_e, H_{ee}, H_{ei}, H_i , and H_{ii} denote contributions from kinetic energy of electrons, electron-electron interactions, electron-ion interactions, kinetic energy of ions and ion-ion interactions, respectively. In equation (11) $\{R_i\}$ and $\{r_l\}$ are electronic and ionic coordination; $\{p_i\}$ and $\{P_l\}$ electronic and ionic momenta, and, M and m are corresponding masses. v and w denote electron-ion and ion-ion potential energies, respectively.

In the canonical ensemble theory, the normalized equilibrium probability density $f_0^{(N)}$ for a system of homonuclear atoms is given by

$$f_0^{(N)}(\vec{r}_1, \dots, \vec{r}_N, \vec{p}_1, \dots, \vec{p}_N) = \frac{\exp[-\beta H(\vec{r}_1, \dots, \vec{r}_N, \vec{p}_1, \dots, \vec{p}_N)]}{N! h^{3N} Q_N(V, T)} \quad (12)$$

where h denotes Planck's constant, and $Q_N(V, T)$ the total partition function,

$$\begin{aligned} Q_N(V, T) &= \text{Tr} e^{-\beta H} \\ &= \frac{1}{N! h^{3N}} \int d\vec{r}_1 \dots d\vec{r}_N \int d\vec{p}_1 \dots d\vec{p}_N \text{Tr}_e e^{-\beta H} \end{aligned} \quad (13)$$

where Tr_e refers complete set of electronic states corresponding to a particular ionic configuration. The motion of ions is very slow relative to the conduction electrons, so, ions can be treated classically unlike electrons that must be handled quantum mechanically. As classical particle do not

obey uncertainty principle one can integrate over position and momentum independently. The result thus obtained is

$$Q_N(V, T) = \frac{1}{N!} \left[\frac{2\pi m}{\hbar^2 \beta} \right]^{\frac{3N}{2}} Z_N(V, T) \quad (14)$$

where the configurational partition function

$$Z_N(V, T) = \int \cdots \int d\vec{r}_1 \cdots d\vec{r}_N \exp(-\beta H_{ii}) \left\{ \int \cdots \int d\vec{R}_1 \cdots d\vec{R}_N d\vec{P}_1 \cdots d\vec{P}_N e^{-\beta(H_e + H_{ce} + H_{ei})} \right\} \quad (15)$$

Ions move in the following effective pair potential

$$U_N = H_{ii} + F' \quad (16)$$

where F' is the Helmholtz free energy of the conduction electrons in the external potential H_{ei} . F' can be calculated by some approximation schemes. Therefore

$$Z_N = \int d\vec{r}_1 d\vec{r}_2 \cdots d\vec{r}_N e^{-\beta U_N} \quad (17)$$

and the L -body probability density

$$n_N^{(L)} = \frac{\int \cdots \int d\vec{r}_{L+1} d\vec{r}_{L+2} \cdots d\vec{r}_N e^{-\beta U_N}}{Z_N} \quad (18)$$

This is related to the L -particle distribution function defined as

$$\begin{aligned} g^{(L)}(\vec{r}_1, \vec{r}_2, \cdots, \vec{r}_N) &\equiv \frac{n_N^{(L)}}{n^L} \\ &= \frac{N!}{n^L (n-L)!} \frac{\int \cdots \int d\vec{r}_{L+1} \cdots d\vec{r}_N e^{-\beta U_N}}{Z_N} \end{aligned} \quad (19)$$

Now, the two-body reduced distribution function stands as

$$g^{(2)}(\vec{r}_1, \vec{r}_2) = \frac{N(N-1)}{n^2} \frac{\int \cdots \int d\vec{r}_3 \cdots d\vec{r}_N e^{-\beta U_N}}{Z_N} \quad (20)$$

For an isotropic liquid

$$g^{(2)}(\vec{r}_1, \vec{r}_2) = g(|\vec{r}_2 - \vec{r}_1|) = g(r)$$

which is also known as pair correlation function and is the central idea in most liquid state theories.

Now if it is assumed that the effective interionic potential is pairwise additive in the following way

$$U_N = NE(V) + \frac{1}{2} \sum_{i,j} v(r_{ij}) \quad (21)$$

where $E(V)$ is the volume dependent (but structure independent) part of energy that includes the free energy of electrons, then all thermodynamic functions can be expressed in terms of $g(r)$ and

the pairpotential of interaction. We note here that equation (20) cannot be solved analytically even if Eqn. (21) is used. For this mathematical limitation different approximation methods and computer simulation methods are devised to solve for $g(r)$. But for the hard sphere (HS) potential

$$v_{hs}(r) = \begin{cases} \infty & \text{for } r < \sigma \\ 0 & \text{for } r > \sigma \end{cases} \quad (22)$$

$g(r)$ can be evaluated analytically [33,40], here σ denotes the hard sphere diameter (HSD). The pair correlation function for a liquid binary alloy [50] may be expressed as

$$g_{ij}(r) = 1 + \frac{1}{(2\pi)^3 \rho_i \rho_j \sqrt{C_i C_j}} \int (S_{ij}(q) - \delta_{ij}) e^{i\vec{q}\cdot\vec{r}} d^3q \quad (23)$$

where $S_{ij}(q)$ is the static structure factors and q the momentum transfer.

3.2 Thermodynamic perturbation theory

The thermodynamic perturbation theory proposed by Weeks-Chandler-Andersen (WCA) [51] splits the interionic potential as core and tails terms

$$v(r) = v_{\text{core}}(r) + v_{\text{tail}}(r) \quad (24)$$

The core term is related to the HS potential, v_{hs} through the Mayer's cluster expansion in the following way.

$$f_{\mu}(r) = f_{hs}(r) + \mu \Delta f(r) \text{ for } 0 \leq \mu \leq 1 \quad (25)$$

where μ is the coupling parameter, and

$$\Delta f(r) = f_{\text{core}}(r) - f_{hs} = [e^{-\beta v_{\text{core}}} - e^{-\beta v_{hs}}] \quad (26)$$

The Helmholtz free energy can be expanded now as

$$F_{\text{core}} = F_{hs} + E(V) - \frac{1}{2} kT \rho \sigma \xi + \mathcal{O}(\xi^4) \quad (27)$$

where

$$\xi = \frac{1}{\sigma} \int_0^{\infty} B_{hs}(r) d\vec{r} \quad (28)$$

with the blip function

$$B(r) = y_{\sigma}(r) \{e^{-\beta v_{\text{core}}(r)} - e^{-\beta v_{hs}(r)}\}. \quad (29)$$

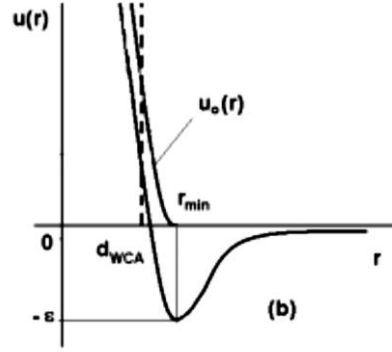


Figure 1: Splitting of the effective pair potential into hard and soft parts.

From equation (27) it is clear that, when $\xi = 0$, $F_{\text{core}} = F_{\text{hs}} + E(V)$. In the WCA theory hard sphere diameter σ is determined following this condition that Fourier transform of $B(r)$ that is $B(q)$ vanishes at $r = \sigma$. But in WCA theory $r^2 B(r)$ shows a saw tooth shaped function. If this function is linearized to have a triangular form one can find an equation [52]

$$\beta v(\sigma) = \ln \left(\frac{-2\beta v'(\sigma) + X + 2}{-\beta v'(\sigma) + X + 2} \right), \quad (30)$$

here, prime indicates the first derivative of the potential energy at $r = \sigma$, and

$$X = \frac{\sigma/\sigma_w}{g_0} \left[\sum_{k=0} \frac{\xi_{k+1}(\eta_w)}{n!} \left(\frac{\sigma}{\sigma_w} - 1 \right)^n - \frac{A\sigma_w}{\sigma^2} (1 + \mu\sigma) \right] \quad (31)$$

All symbols are defined in reference [52]. Solution of the transcendental equation (30) yields the effective HSD. The pair correlation function is now evaluated using this effective HSD.

Andersen et al. [53] proposed a simplest version of the perturbative scheme

known as exponential approximation,

$$g_{\text{hs}}(r) = g e^{-v(r)/kT}$$

where $v(r)$ is the real short-range part of the potential; in the present case it is v_{core} . We note that this optimized form gives more realistic description of the pair correlation function.

Now using the perturbation theory one can calculate the free energy of a system per ion as

$$F = F_{\text{unp}} + 2\pi\rho \int v_{\text{pert}} g_{\text{hs}} d^3r \quad (32)$$

where

$$F_{\text{unp}} = E(V) + F_{\text{hs}} = F_{\text{vol}} + F_{\text{gas}} + F_{\text{hs}}$$

and

$$F_{vol} = \frac{1}{32\pi^3} \int_0^{\infty} q^4 \left\{ \frac{1}{\epsilon(q)} - 1 \right\} |v(q)|^2 dq - \frac{ZE_F}{3Y} \quad (33)$$

where $Z = C_1Z_1 + C_2Z_2$ $Y = \chi_{elec}/\chi_F$, subscripts elec and F denote isothermal compressibility of the interacting and free electrons, respectively. The values of Y are obtained from [54].

The electron gas contribution to the free energy per valence in Rydberg unit is

$$F_{gas} = \frac{2.21}{r_s^2} - \frac{0.916}{r_s} + 0.31 \ln r_s - 0.115 \quad (34)$$

where

$$\begin{aligned} r_s &= \left(\frac{3}{4\pi\rho Z} \right)^{\frac{1}{3}} / a_0; \quad \rho = \frac{\rho_1\rho_2}{C_1\rho_2 + C_2\rho_1} \\ F_{hs} &= \sum_i [-\ln(\Lambda_i^3 v) + \ln C_i] - \frac{2}{3} \left(\frac{5}{3} - y_1 + y_2 + y_3 \right) \\ &\quad + (3y_2 - 2y_3)/(1 - \eta) + \frac{3}{2} \left(1 - y_1 - y_2 - \frac{y_3}{3} \right) / (1 - \eta)^2 \\ &\quad + (y_3 - 1) \ln(1 - \eta) \end{aligned} \quad (35)$$

where

$$\begin{aligned} \Lambda_i &= \left\{ \frac{2\pi\hbar^2}{m_1^{c_1} m_2^{c_2} kT} \right\}^{\frac{1}{2}} \\ \eta &= \sum_i \eta_i; \quad \eta_i = \frac{C_i \pi \rho_i \sigma_{ii}^3}{6} \\ F_{tail} &= D \sum_{i,j} C_i C_j M_{ij} \end{aligned}$$

$$D = 2\pi\rho,$$

$$M_{ij} = \int_{\sigma}^{\infty} v_{ij} g_{ij}(r) r^2 dr$$

Now, the energy of mixing

$$\Delta F = \Delta F_{vol} + \Delta F_{gas} + \Delta F_{hs} + \Delta F_{tail} \quad (36)$$

ΔF_y to be calculated by using equation (8). Now if the experimental densities of the alloy at different concentrations are available, and if the difference between calculated density and the experimental ones exists and significant an excess volume correction to be added with the thermodynamics of mixing as [34]

$$\Delta F = \Delta F_{vol} + \Delta F_{gas} + \Delta F_{hs} + \Delta F_{tail} + \Delta F_{evc}. \quad (37)$$

Enthalpy of alloy:

Enthalpy of the alloy per ion

$$\begin{aligned}
H &= E + pV \\
&= \frac{3}{2}kT + E(V) + \frac{\rho}{2} \sum_{i=1}^2 \int g_{ij}(r)v_{ij}(r)d^3r + pV
\end{aligned} \tag{38}$$

Entropy of alloy:Within the above perturbation scheme the entropy of alloy (divided by Nk) reads [38]

$$\begin{aligned}
S &= S_{\text{ref}} + S_{\text{tail}} \\
S_{\text{ref}} &= S_{\text{id}} + S_{\text{gas}} + S_{\eta} + S_{\sigma} \\
S_{\text{id}} &= -[C_1 \ln C_1 + C_2 \ln C_2] \\
S_{\text{gas}} &= \frac{5}{2} \ln \left[\frac{1}{\rho} \left(\frac{m_1^{C_1} m_2^{C_2} kT}{2\pi\hbar^2} \right)^{\frac{3}{2}} \right] \\
S_{\eta} &= \ln(1-\eta) + \frac{3}{2}[1 - (1-\eta)^{-2}] \\
S_{\sigma} &= \left[\frac{\pi C_1 C_2 \rho (\sigma_{11}^2 - \sigma_{22}^2) (1-\eta)^{-2}}{24} \right] \\
&\quad \times \{12(\sigma_{11} + \sigma_{22}) - \pi\rho[C_1\sigma_{11}^4 + C_2\sigma_{22}^4]\}
\end{aligned} \tag{39}$$

and

$$S_{\text{tail}} = \frac{1}{k} \left[\left(\frac{\partial F_{\text{tail}}}{\partial T} \right)_{v,\rho,\sigma_{ii}} + \sum_{i=1}^2 \left(\frac{\partial F_{\text{tail}}}{\partial \sigma_{ii}} \right)_{v,T} \left(\frac{\partial \sigma_{ii}}{\partial T} \right)_{v,\rho} \right] \tag{40}$$

The temperature dependent HSD as proposed by Protopapas et al. [55] is

$$\sigma(T) = 1.126\sigma_m \left\{ 1 - 0.112 \left(\frac{T}{T_m} \right)^{\frac{1}{2}} \right\}. \tag{41}$$

Entropy of mixing therefore stands

$$\Delta S = \Delta S_{\text{ref}} + \Delta S_{\text{tail}} \tag{42}$$

An alternative way may also be used to evaluate entropy of mixing

$$\Delta S = \frac{\Delta H - \Delta F}{T} \tag{43}$$

3.3 The pseudopotential model

The effective electron-ion interaction between a conduction electron and an ion may be written as (in atomic unit) [56]

$$w(r) = \begin{cases} \sum_{m=1}^2 B_m \exp(-r/ma) & \text{if } r < R_c \\ -Z/r & \text{if } r > R_c \end{cases} \quad (44)$$

where Z, R_c and a are the effective s -electron occupancy number, core radius and, the softness parameter, respectively. B_m is the coefficient of expansion which is independent of r but depends explicitly on parameters Z, R_c and a . The pseudopotential theory leads to an expression for effective interionic potential of an alloy through the energy band structure [33,36],

$$v_{ij}(r) = \frac{Z_i Z_j}{r} \left[1 - \frac{2}{\pi} \int dq F_{ij}^{(N)} \frac{\sin qr}{q} \right] \quad (45)$$

where the wave number characteristics

$$F_{ij}^{(N)} = \left[\frac{q^2}{8\pi\rho\sqrt{Z_i Z_j}} \right]^2 w_i(q) w_j(q) \left[1 - \frac{1}{\epsilon(q)} \right] [1 - G(q)]^{-1} \quad (46)$$

3.4 Noticeable beckon and phenomenological theory of segregation

For a condensed state one of the most basic ingredients from which any microscopic description begins is the subatomic interaction or interionic interaction derived from the former one. This interaction dictates if the alloy would be an ordered or a segregating type. In ordered alloy, the unlike atoms are preferred as nearest neighbors to like atoms, whereas in segregating alloys like atoms are preferred as nearest neighbors to unlike atoms. But direct identification of like and unlike atoms in the sample is very difficult to achieve experimentally. Indirect ways through some probes assigned with interionic interactions, for example, structural data, thermodynamics of mixing (viz. energy of mixing, enthalpy of mixing, entropy of mixing), atomic transport properties (viz. coefficient of shear viscosity, diffusion coefficient) and electronic transport properties (viz. resistivity) provide good alternative ways. Some of the microscopic parameters used in identifying segregating alloys are

- (i) downward concavity or positivity of the free energy of mixing vs concentration profile at any or some concentrations,
- (ii) upward concavity or negativity of the enthalpy of mixing profile at any or some concentrations,
- (iii) upward concavity or negativity of the entropy of mixing profile at any or some concentrations,
- (iv) order potential $v_{\text{ord}} = v_{ij}(r) - \frac{v_{ii} + v_{jj}}{2} > 0$ around the nearest neighbour distance,
- (v) a strong bending of the viscosity vs concentration curve near the critical concentration at and below the critical temperature.

In the phenomenological theories or empirical methods [26] there are some other parameters also to identify the segregation of alloys, for example,

- (i) the Warren-Cowly short range order parameter $\alpha > 0$,
- (ii) in the regular solution theory, the exchange energy $w > 0$,
- (iii) the concentration-concentration structure factors in the long wavelength limit $S_{cc}(0)$ diverges near the critical temperature and, the sharp increase happens around the critical concentration.

Some alloys such as Li – Na, Al – Bi, Al – Sn, Fe – Cu, Cu – Co, Al – Pb, Bi – Zn, Cd – Ga, Ga-Pb, Ga-Hg, Pb-Zn, Pb-Si, and Cu-Pb are well known systems for which some segregating properties are measured. So, it is worth pursuing to understand the microscopic origin of segregation from the theoretical point of view and compare them with the experimental ones.

3.5 Phenomenological theories in the study of segregating properties

The Gibbs free energy of mixing (of a sample of N moles) for binary alloys is

$$\Delta G = G_{\text{alloy}} - \sum_{i=1}^2 C_i G_i. \quad (47)$$

In terms of the partial Gibbs energies ΔG_i , one can write

$$\Delta G = C_i \Delta G_i + C_j \Delta G_j, \quad (48)$$

with

$$\Delta G_i = RT \ln a_i \quad (i = 1,2)$$

where a_i denotes the thermodynamic activity of the i -th component.

The stability of a binary mixture is mostly determined by ΔG . Figure 2(i) shows a schematic diagram for ΔG denoted by G_M as a function of concentration C . Here curve a describes a miscible stable state whereas curve b describes an immiscible unstable state in the concentration range ΔC . The points P and Q in Figure 2(i) give compositions of two segregated phases. At points P and Q the partial Gibbs energies of the components are equal,

$$\Delta G_i(C_1) = \Delta G_i(C_2) \quad (i = A, B).$$

The point of inflexion in the curve b for $T_2 < T_c$ represents the spinodal line. The critical concentration and critical temperature follow from the following conditions

$$\left(\frac{\partial^2 \Delta G}{\partial C^2} \right)_{C=x_c} = 0; \quad \left(\frac{\partial^3 \Delta G}{\partial C^3} \right)_{C=x_c} = 0$$

at $T = T_c$. Following the Bhatia-Thronston structure factors [57], which is well known for the concentration-concentration fluctuation in the long wave length limit, one can show

$$S_{CC}(0) = RT \left(\frac{\partial^2 G_m}{\partial C^2} \right)_{T,p}^{-1}.$$

As

$$C \rightarrow x_c, \quad \text{and} \quad T \rightarrow T_c, \quad S_{CC}(0) \rightarrow \infty$$

This property of $S_{CC}(0) \rightarrow \infty$ signals the phase separation in a binary mixture. Figure 2(ii) shows this behaviour. Other empirical models used

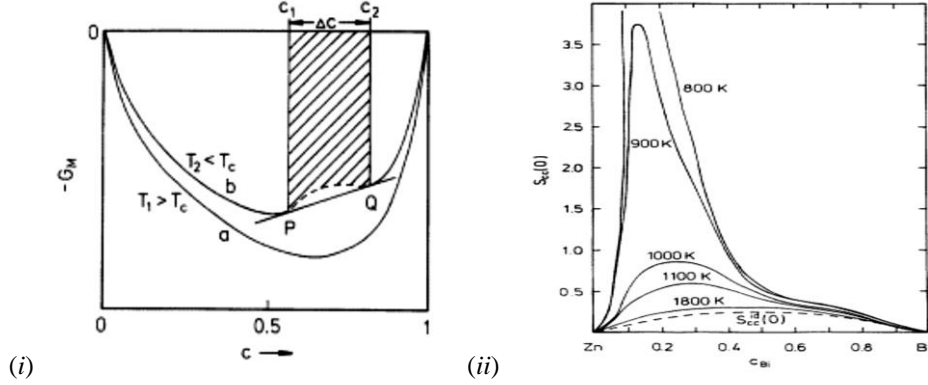


Figure 2: (i) A schematic diagram of Gibbs free energy of mixing as a function of concentration and (ii) $S_{cc}(0)$ for different temperatures (after Singh and Sommer [26]).

Table 1: Critical concentration and critical for demixing liquid alloys.

Systems				
$A_m - B_n$	m	n	$x_{c,A}$	$\frac{w}{kT_c}$
A-B	1	1	0.5	2.0
A ₂ -B ₂	2	2	0.5	1.0
A ₄ -B ₄	4	4	0.5	0.5
A-B ₂	1	2	0.74	1.457
A ₂ -B ₄	2	4	0.74	0.728
A-B ₃	1	3	0.84	1.244
A ₂ -B ₈	2	8	0.89	0.562

in the study of demixing of alloys are quasi-lattice theory [7,26] and the self association model [58]. Using the quasi-lattice theory [7,26] it is possible to derive the configurational energy and partition function of the alloy. This knowledge later yields an expression for the Gibbs free energy of mixing and thermodynamic activity in terms of a free parameter known as interchange energy. The critical properties of segregation can then be obtained from the so-called stability conditions

$$\frac{\partial \ln a_i}{\partial C_i} = 0; \quad \frac{\partial^2 \ln a_i}{\partial C_i^2} = 0$$

for different clusters suggested by the self-association model.

Although this empirical theory presents a good prescription to study the critical properties of segregating alloys, its reliability in predicting critical properties of real binary alloys is yet to be seen.

4. RESULTS FROM THE MICROSCOPIC APPROACH

4.1 Partial pair potentials and corresponding pair correlation functions

For any microscopic description of a condensed matter the most fundamental ingredient necessary is the knowledge of interionic potential. Direct derivation and application of the N -body potentials to the study of the physical properties of condensed matter is a too much difficult job to handle theoretically. In order to avoid this difficult situation, one goes for the effective pair potentials. The term effective indicates that, these potentials take into account the many body effects in an average way following indirect routes. Figure 3 shows the profile of the effective partial pair potentials for an Al-based alloy namely $\text{Al}_x\text{In}_{1-x}$. It is seen that partial pair potential $v_{\text{AlAl}}(r)$ has the shallowest potential well and v_{InIn} the deepest well. That of v_{AlIn} lies in between. It is also seen that the position of the well minima for v_{AlIn} and v_{InIn} shift to large r relative to v_{AlAl} . Similar feature is also observed for transition metal segregation alloys (for example $\text{Fe}_x\text{Cu}_{1-x}$, $\text{Cu}_x\text{CO}_{1-x}$). In case of $\text{Zn}_x\text{Bi}_{1-x}$ the amount of shift among different partial pair potentials is significantly small. This shifting is largely associated with the difference in the values of the core radii between individual components of the alloy. We note here that in random alloys v_{12} generally lies between v_{11} and v_{22} . But in the case of compound forming alloys v_{12} goes down the well of the v_{11} or v_{22} whichever has lower value. We note that in the effective pair potential calculations Ichimaru-Utshumi dielectric function [59] has been used by Bhuiyan and his group because this function satisfies both compressibility sum rule and the short range correlation function. The BS pseudopotential model has proven to be successful in the studies of liquid structure [60-63], thermodynamic properties [34,35,64,65], atomic transport [66-70] and electronic transport properties [71,72] of liquid metals and there alloys.

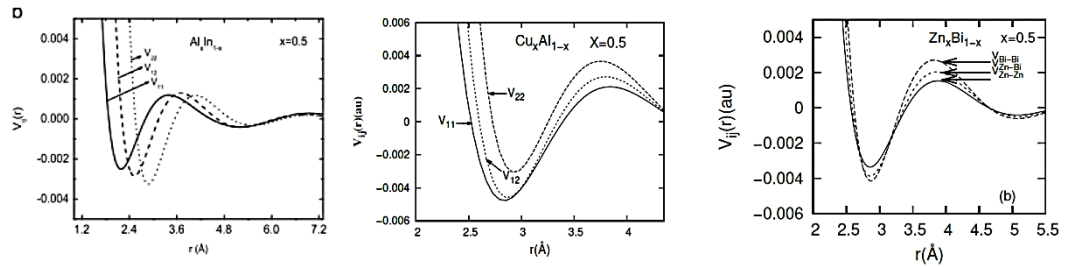


Figure 3: Partial pair potentials for $\text{Al}_x\text{In}_{1-x}$, $\text{Cu}_x\text{Al}_{1-x}$ and $\text{Bi}_x\text{Zn}_{1-x}$ liquid binary alloys For $x = 0.5$ (after (from the left) Bhuiyan and coworkers [21, 22, 24]).

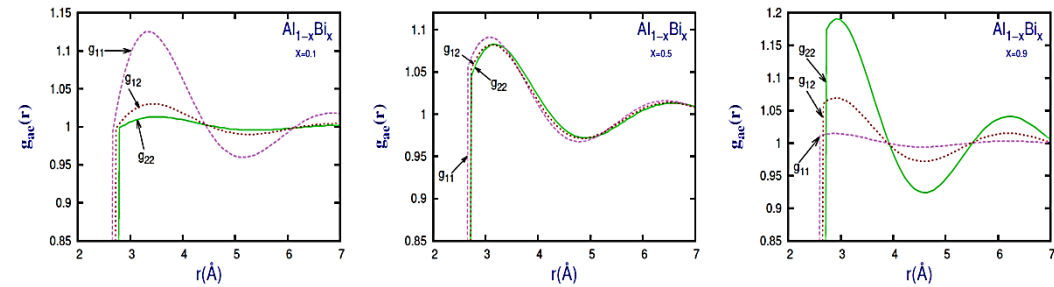


Figure 4: Partial pair potentials for $\text{Bi}_x\text{Zn}_{1-x}$ liquid binary alloys for $x = 0.1$, $x = 0.5$, $x = 0.9$, respectively (after Kasem et al. [24]).

The partial pair correlation function, $g(r)$, is related to the partial interionic pair potential through the statistical mechanics [73] (see equation (7)). Partial pair correlation functions for three different concentrations are presented in figure 4. In the alloys, rich in component 1, g_{11} exhibits the largest peak, while the trends become opposite in alloys rich in component 2; that is g_{22} shows the largest main peak. But in both cases peak value of g_{12} remains in the middle of $g_{11}(r)$ and $g_{22}(r)$. The physical significance of $g(r)$ is that, it gives a measure of the probability of finding the number of nearest neighbors at a distance of the peak from the ion located at the origin. Thus, the area under the principal oscillation provides the coordination number, a characteristic feature of the condensed matter. Advantage of it is that, $g(r)$ can also be derived from the X-ray or neutron diffraction data through the Fourier transformation, and directly from the computer simulation experiment. In the theoretical study of liquid metals, it plays the central role in describing thermodynamic properties.

4.2 Energy of mixing

The free energy of mixing and its effects on the critical properties of segregation are described for different alloys below.

(a) Li_xNa_{1-x} liquid binary alloys:

The first attempt to estimate the energy of mixing theoretically for $Li_{1-x}Na_x$ liquid binary alloys from a microscopic approach was made by Tamaki [74]. He also attempted to relate effective pair potential between ions with the immiscibility of the segregating alloys (see Figure 5(a)). He was Stroud who made an attempt systematically for the first time to understand the segregating properties such as critical concentration x_c and critical temperature T_c of $Li_{1-x}Na_x$ liquid binary alloys using a microscopic theoretical approach [23]. He employed there the electronic theory of metal based on the empty core model [75], statistical mechanics and the Gibbs-Bogoliubov variational scheme [76] in order to calculate the free energy of mixing. Figure 5(b) shows a schematic diagram, presented by Stroud in [23], for the ΔF as a function of concentration for different temperatures. For $T > T_c$ the energy of mixing profile is concave upward for all concentrations, which manifests complete miscibility (i.e., alloy is stable against segregation) at any concentration. But for $T < T_c$ the profile becomes concave downward at some concentrations which indicates segregation of the alloy. The temperature at which spinodal points P and Q coincides is called the critical temperature T_c , and the concentration at which it happens is called the critical concentration x_c . Here, in the calculation the Hubbard type dielectric function [77] is used. The critical concentration for Li_xNa_{1-x} segregating alloy was found to be $x_c = 0.7$, but the predicted critical temperature was overestimated by one third [23].

Table 2: Potential parameters and densities used for elements that formed different alloys under study are listed.

Elements	ρ ($^{\circ}A^{-3}$)	R_c (au)	a (au)	Z
Al	0.0517	1.91	0.30	3.0
In	0.0342	1.32	0.29	3.0
Bi	0.0289	1.49	0.36 (0.35)	3 (5)
Fe	0.0756	1.425	0.33	1.5
Co	0.0787	1.325	0.27	1.5
Cu	0.0760	1.510	0.44	1.5

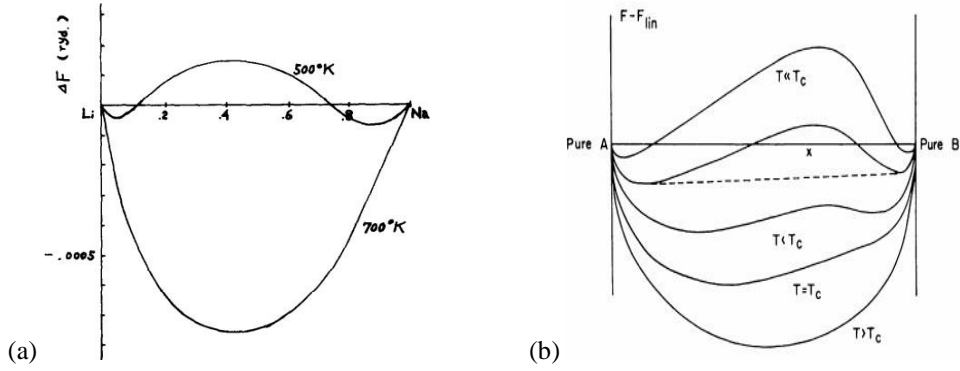


Figure 5: Energy of mixing as a function of x for $\text{Li}_x\text{Na}_{1-x}$ liquid binary alloys (after (a) Tamaki [72], (b) Stroud [20])

(b) $\text{Al}_x\text{In}_{1-x}$ liquid binary alloys:

This $\text{Al}_x\text{In}_{1-x}$ alloy is formed by the elemental metals Al and In. These elements belong to the less simple polyvalent metals. Al based alloys are known to be good candidates for a new advanced anti-friction material. The input values such as potential parameters R_c , a and Z along with number density, ρ , for Al and In, and also for some other elements are shown in Table 2.

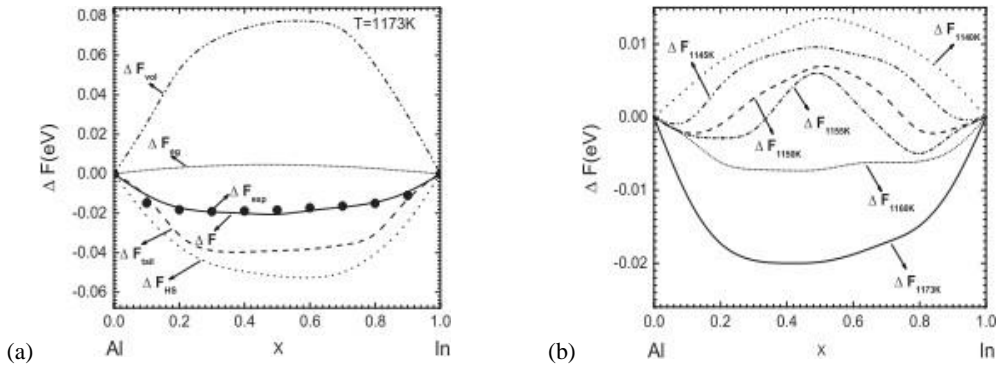


Figure 6: Energy of mixing as a function of x for $\text{Al}_x\text{In}_{1-x}$ liquid binary alloys (a) breakdown details at $T=1173\text{K}$, (b) Temperature dependence (after Faruk and Bhuiyan [21]).

Faruk and Bhuiyan [21] studied the segregating properties of $\text{Al}_x\text{In}_{1-x}$ liquid binary alloys by using the electronic theory of metals (first principal calculations) along with the statistical mechanics and perturbative approach. Initially, they justified the appropriateness of the potential parameters by calculating static structure factors of the elemental liquids at a thermodynamic state at which experimental data are available [78]. Figure 6(a) shows the breakdown details of energy of mixing, ΔF , at $T = 1173\text{K}$. It is noticed that the HS contribution to the energy of mixing is negative and values are the lowest among all other contributions across the whole range of concentrations. The tail part contribution is also negative across the concentration range and values are the second lowest among all others. Contribution of the electron gas, ΔF_{eg} , is positive for the full concentration range and values are very close to zero. The volume dependent (i.e., structure independent) part of the energy of mixing, ΔF_{vol} , due to electron ion interaction is positive and large across the full range of

concentration. The combined effect of all contributions, ΔF , agree well with the corresponding experimental data [79]. This signifies the accuracy of the approach for the study of energy of mixing at different temperatures.

The temperature dependent energy of mixing for $\text{Al}_x\text{In}_{1-x}$ liquid binary alloys for different concentrations are illustrated in figure 6(b). As temperature is decreased from 1173 K, ΔF increases gradually and at 1155 K becomes partially positive and partially negative. Further lowering of temperature increase the miscibility gap and at 1140 K the concentration gap span the whole range of concentration. A careful observation finds the first downward concavity or positivity of ΔF at 1160 K, and the concentration at which it happens is $x = 0.5$. So, the predicted critical temperature and concentration for $\text{Al}_x\text{In}_{1-x}$ segregating alloys are $T_c = 1160$ K and $x_c = 0.5$, respectively. The experimental work by Campbell et al. [80], and Campbell and Wagemann [9] report a critical temperature of 1220 K. whereas Predel [1] reports 1100 K for $\text{Al}_x\text{In}_{1-x}$ liquid binary alloys. Differential thermal analysis by Sommer et al. [81] reports $T_c = 1112$ K. The average of these scattered experimental data is 1144 K which is close to the theoretical prediction of Faruk and Bhuiyan [21]. The experimental critical concentration [10,1] is $x_c = 0.5$ which is exactly the same as that of theoretical prediction [21]. But the experimental data reported in [9] is 0.34 which largely deviates from 0.5.

(c) $\text{Bi}_x\text{Al}_{1-x}$ liquid binary alloys:

$\text{Bi}_x\text{Al}_{1-x}$ liquid binary alloy is formed by two elements Al and Bi which belong to group IIIB and VB in the periodic table, respectively. The melting points of Al and Bi are 933 and 544 K, respectively; the corresponding densities are 2.375 and 9.78 gm cm^{-3} . Al is a trivalent and Bi is a pentavalent metal. The atomic radii of Al and Bi are 1.82 and 1.63 Å, respectively. The large mismatch in their physical properties makes this alloy interesting to study theoretically.

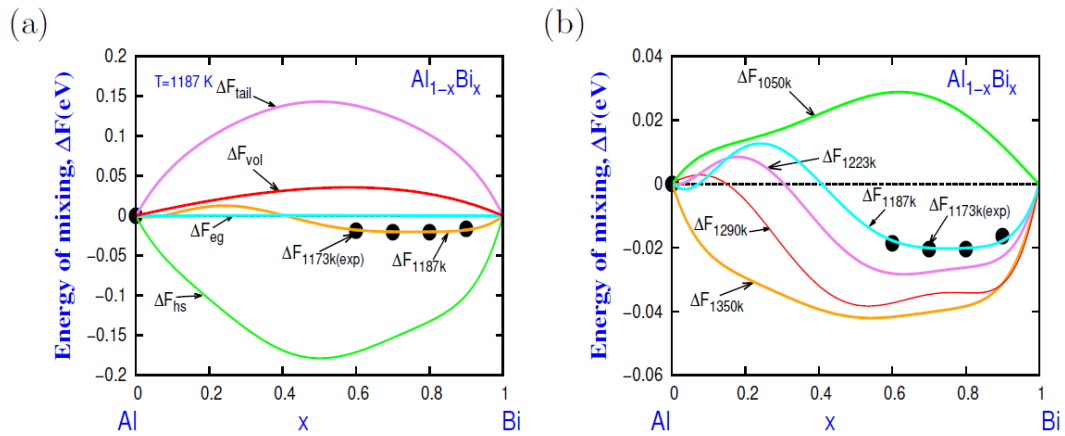


Figure 7: Energy of mixing as a function of x for $\text{Bi}_x\text{Al}_{1-x}$ liquid binary alloys (a) breakdown details at $T = 1187$ K, (b) Temperature dependence (after Abbas et al. [25]).

Figure 7(a) illustrates the breakdown details of the energy of mixing at $T = 1187$ K at which some experimental data [79] for ΔF are available in the literature. The HS contribution to the energy of mixing is negative for the whole concentration range as is found for $\text{Al}_x\text{In}_{1-x}$ liquid binary alloys. But unlike $\text{Al}_x\text{In}_{1-x}$, the tail part contribution, ΔF_{tail} , of $\text{Bi}_x\text{Al}_{1-x}$ alloys is positive for all

concentrations with a maximum near equiatomic concentration. The volume dependent part, ΔF_{vol} , in this case, is positive but the magnitudes are much lower than that of ΔF_{tail} . The electron gas contribution, ΔF_{eg} , is nearly zero as for Al_xIn_{1-x} . The combined effect of all contributions to the free energy, however, agrees well with the experimental results at $T = 1173$ K available in the literature [79].

Figure 7(b) shows the energy of mixing for Bi_xAl_{1-x} liquid binary alloys for different temperatures. It appears that the alloy exhibits a complete miscibility at 1350 K, and immiscibility for all concentrations at 1050 K. But at $T = 1290$ K, ΔF shows a partial positivity with concavity downward near $x = 0.15$. Further decrease of temperature gradually enhances the concentration gap. As the concavity downward (or positivity) of ΔF manifests onset of segregation, one can conclude that the predicted critical concentration is $x_c = 0.15$, and critical temperature $T_c = 1290$ K, while the corresponding experimental values are $x_c = 0.19$ [82] and $T_c = 1310$ [82,83].

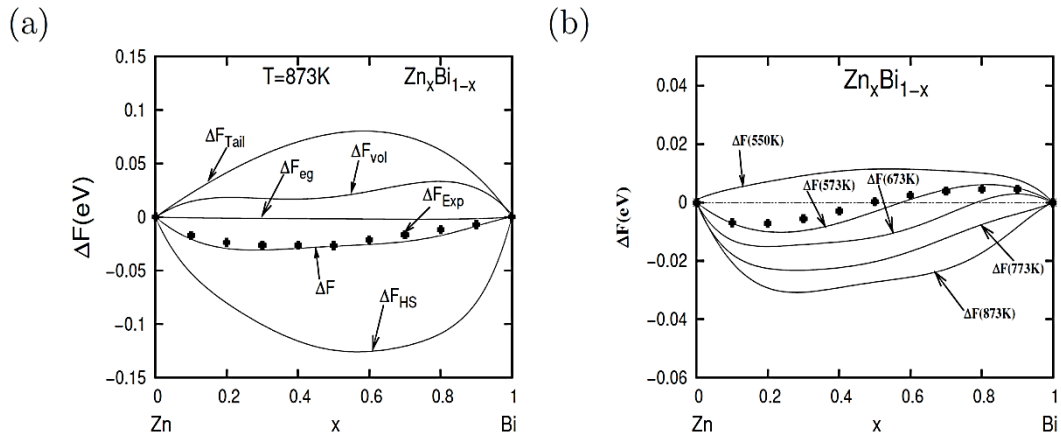


Figure 8: Energy of mixing as a function of x for Bi_xZn_{1-x} liquid binary alloys (a) breakdown details at $T=873$ K, (b) Temperature dependence (after Kasem et al. [24]).

(d) Zn_xBi_{1-x} liquid binary alloys:

Figure 8(a) shows the free energy of mixing for liquid Zn_xBi_{1-x} alloys at 873 K [24]. ΔF_{hs} is negative for Zn_xBi_{1-x} liquid binary alloys for the whole range of concentration, this trend is similar to that of the previous alloys, and having the smallest values relative to the other components for each concentration. In this case ΔF is asymmetric in nature where the minimum value is found to be around $x = 0.6$ which is located in the Bi rich alloys. ΔF_{tail} and ΔF_{vol} contributions are positive for the full range of concentrations, but ΔF_{tail} shows the larger values than that of ΔF_{vol} . ΔF_{eg} contribution is nearly zero as other segregating alloys under consideration of [21] this review article. The total energy of mixing, however, matches well with corresponding experimental data [79].

Temperature dependence of ΔF are illustrated in figure 8(b) It is noticed that at $T=773$ K and higher temperatures ΔF is negative for all concentrations. This nature indicates that the alloy is completely miscible in the regime of the above thermodynamic states. But at a lower temperature $T = 673$ K, ΔF becomes positive i.e., concave downward for some concentrations and negative for others. When temperature is lowered further miscibility gap increases gradually as previous systems and cover the whole concentration range at 550 K. From figure it appears that the critical concentration is $x_c = 0.9$

and the critical temperature $T_c = 773$ K. The experimental value for x_c is 0.83 [16,17], the critical concentration found theoretically by Stroud [36] and Karlhauber et al [84] (from quasi lattice theory) was $x_c=0.75$ and 0.87, respectively. Experimental critical temperatures are 856 K [16] and 878 K [17], and a theoretical study shows 438 K [36].

(e) Fe_xCu_{1-x} liquid binary alloys:

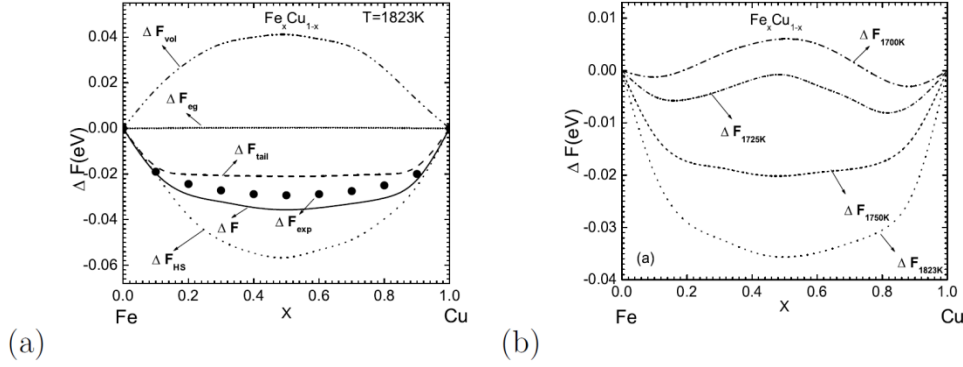


Figure 9: Energy of mixing as a function of x for Fe_xCu_{1-x} liquid binary alloys (a) breakdown details at $T=1823$ K, (b) Temperature dependence (After Faruk et al. [22]).

Figure 9(a) shows that the HS contribution to the free energy of mixing for Fe_xCu_{1-x} is negative for all concentrations at $T = 1823$ K [22]. Here the tail part contribution to the energy of mixing is negative for all concentrations unlike other segregating alloys. The volume dependent term ΔF_{vol} is positive for the whole concentration range, and the electron gas contribution ΔF_{eg} is almost zero as for all others discussed above. The total energy of mixing obtained summing all four contributions is negative for all concentrations and the agreement with available measured data [79] is fairly good. At $T = 1823$ K the alloys remain miscible across the full concentration range. As temperature is lowered to 100 K, ΔF becomes partially positive around equiatomic concentration where the concavity is downward, and the other part of energy of mixing remains negative with upward concavity. The critical temperature thus found was $T_c = 1750$ K and the critical concentration found was $x_c = 0.5$. The experimental values reported by different authors for x_c are 0.56 [13], 0.538 [14] and 0.538 [15], and the corresponding experimental data for T_c are 1696, 1704 K and 1694 K, respectively.

(f) Co_xCu_{1-x} liquid binary alloys:

In this case the behavior of various contributions to ΔF is found to be similar to that of Fe_xCu_{1-x} [22]. But, in the immiscible state ΔF shows (Figure 10) an asymmetric feature with a value of critical concentration $x_c = 0.58$ and critical temperature $T_c = 1650$ K. The corresponding experimental values are $x_c = 0.53$ and $T_c = 1547$ K [12].

Bhuiyan and coworkers carefully investigated why ΔF varies with temperature. They have found that ΔF_{hs} and ΔF_{tail} are sensitive to T and are mostly responsible for the variation. While ΔF_{vol} and ΔF_{eg} are not sensitive to T at all. The sensitivity arises, in this case, due to the alteration of σ and consequently $g_{hs}(r)$, with the change of T .

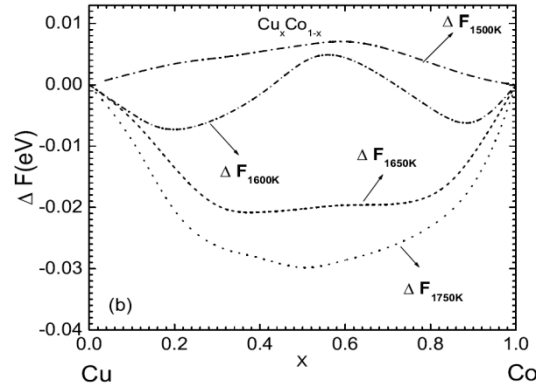


Figure 10: Temperature dependence of energy of mixing as a function of x for $\text{Cu}_x\text{Co}_{1-x}$ liquid binary alloys (after Faruk et al. [22]).

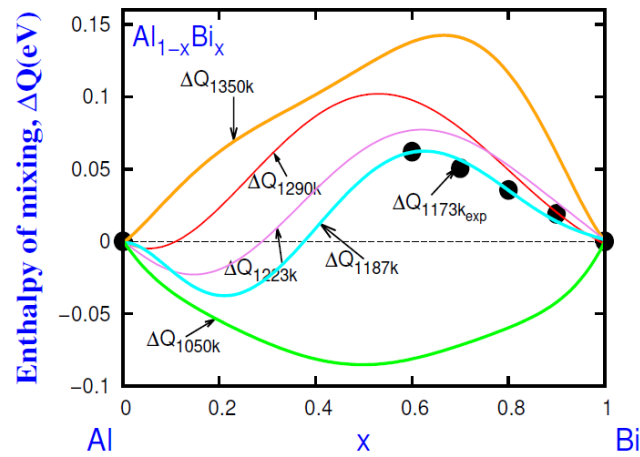


Figure 11: Temperature dependence of enthalpy of mixing as a function of x for $\text{Bi}_x\text{Al}_{1-x}$ liquid binary alloys (after Fysol et al. [25]).

4.3 Enthalpy of mixing

(a) $\text{Bi}_x\text{Al}_{1-x}$ liquid binary alloys:

The enthalpy of mixing, ΔH , are used as a probe to study the critical properties of $\text{Bi}_x\text{Al}_{1-x}$ liquid binary alloys. Figure 11 demonstrates that calculated values of enthalpy of mixing agree in an excellent way with available experimental data for miscible alloys at 1187 K [79]. However, figure also show that the trends of ΔH as a function of concentration is just opposite like a mirror reflection to that of free energy of mixing discussed above. That is at 1350 K ΔH is positive, and at 1050 K it is negative for the full concentration range, while at the same thermodynamic states ΔF shows negative and positive values, respectively. Figure also shows that, at about 1290 K, ΔH exhibits negative (i.e., concave upward) at low values of x and positive (i.e., concave downward) for the rest. That means segregation of the alloy begins at 1290 K which is exactly same as found from ΔF [25].

But in the case of enthalpy of mixing the critical concentration is found to be somewhat smaller than predicted by the energy of mixing [25].

4.4 Entropy of mixing

It is interesting to see how another static magnitude the entropy of mixing describes the critical properties of segregation for different alloys.

(a) Zn_xBi_{1-x} liquid binary alloys

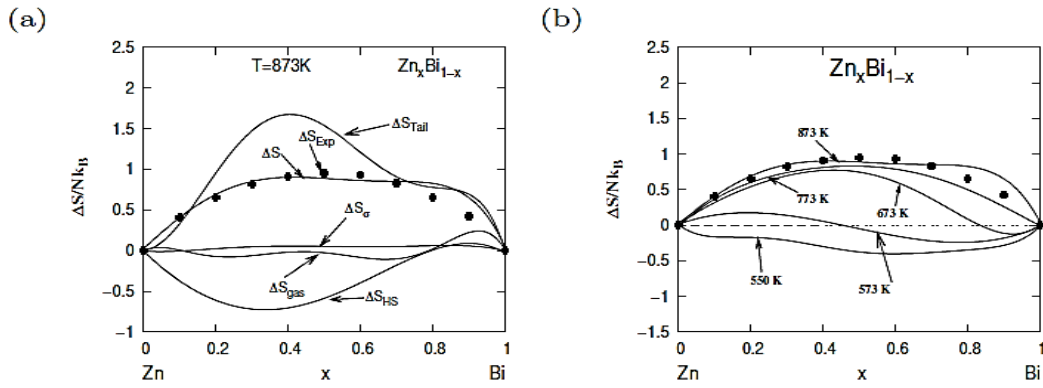


Figure 12: Entropy of mixing as a function of x for Bi_xZn_{1-x} liquid binary alloys; (a) breakdown details, (b) temperature dependence (after Kasem et al. [24]).

Figure 12(a) shows the breakdown details of different contributions to the total entropy of mixing calculated by Kasem et al. [24]. It is seen from figure that at $T = 873 K$, ΔS_{hs} is negative up to $x \leq 0.8$ and then becomes positive. ΔS_{gas} is negative in the concentration interval $0.1 < x < 0.8$, and positive beyond it. Contribution of HSD mismatch term, ΔS_{σ} , is almost zero across the whole range of concentration. The tail part contribution, ΔS_{tail} , is found to be positive for the full concentration range. However, the combined effect of these contributions that is the total entropy of mixing is positive for all concentrations and, the agreement between theory and experiment is very good up to $x = 0.7$, and fairly good for $x > 0.7$ [79].

Figure 12(b) shows the temperature dependence of entropy of mixing for Zn_xBi_{1-x} liquid binary alloys [24]. We note here that negativity of ΔS (i.e., upward concavity) is an indication of segregation. Figure also shows that the critical temperature and critical concentration are $T_c = 773 K$ and $x_c = 0.9$, respectively. These values are found to be same as that found from the energy of mixing [24].

(b) Bi_xAl_{1-x} liquid binary alloys

Figure 13 shows entropy of mixing for Bi_xAl_{1-x} liquid binary alloys calculated by Fysol et al. [25]. At $T = 1350 K$, ΔS is positive for all concentrations and at $T = 1050 K$ it is negative for the whole range of concentration. For temperatures in between the entropy of mixing is partly positive and partly negative. Fysol et al. [25] theoretically found the values $x_c = 0.1$ and $T_c = 1290 K$ for critical concentration and critical temperature, respectively. Here the value of T_c is the same as that found from ΔF [25] but x_c is somewhat lower in this case.

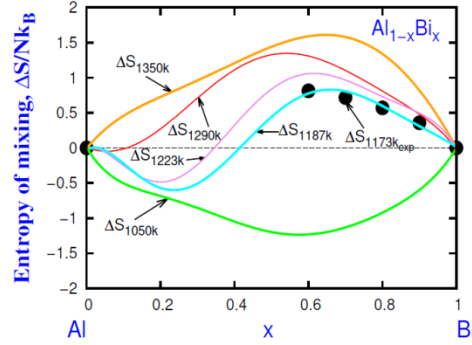


Figure 13: Temperature dependence of entropy of mixing as a function of x for $\text{Bi}_x\text{Al}_{1-x}$ liquid binary alloys (after Fysol et al. [25]).

4. CONCLUDING REMARKS

Looking at figures of free energy of mixing for $\text{Al}_x\text{In}_{1-x}$, $\text{Fe}_x\text{Cu}_{1-x}$, $\text{Cu}_x\text{Co}_{1-x}$, $\text{Zn}_x\text{Bi}_{1-x}$, and $\text{Bi}_x\text{Al}_{1-x}$ liquid binary alloys one can easily find that the HS contribution ΔF_{hs} is always negative for all concentrations and temperatures. This means that, HS liquid alone cannot describe segregation for binary alloys. This finding agrees with that of Libowitz and Rowlinson [40]. However, ΔF_{vol} becomes positive for the whole range of concentration and dominates other contributions in the case of $\text{Fe}_x\text{Cu}_{1-x}$, $\text{Cu}_x\text{Co}_{1-x}$, and $\text{Al}_x\text{In}_{1-x}$ liquid binary alloys; this feature directly favors the segregation for these alloys. The contribution of the tail part of the pair potential, ΔF_{tail} , becomes positive for the full concentration range for $\text{Zn}_x\text{Bi}_{1-x}$, and $\text{Bi}_x\text{Al}_{1-x}$ liquid binary alloys and negative for others. The electron gas contribution ΔF_{eg} is nearly zero for all systems and for any thermodynamic state characterized by temperature. Energy of mixing for hard sphere liquid and the tail part contribution are very sensitive to temperature unlike ΔF_{vol} and ΔF_{eg} . In the case of free energy ΔF_{hs} and ΔF_{tail} increases with increasing temperature, as a result total energy of mixing becomes concave downward which manifests immiscibility of the alloy. The values of the critical temperatures and critical concentrations for different alloys are illustrated in Table 3.

Table: Critical temperature and critical concentrations for different segregating alloys.

Systems	x_c		T_c (K)		Others (Theo.)
	(Theo.)	(Expt.)	(Theo.)	(Expt.)	
AlIn	0.5	0.5, 0.34	1160	1155, 1150, 1145	-
FeCu	0.5	0.56, 0.538	1750	1696, 1704, 1694	-
CuCo	0.58	0.53	1650	15473	-
ZnBi	0.9	0.83	773	856, 878	438
BiAl	0.15	0.19	1290	1310	-

Understanding of the segregating behavior of liquid binary alloys from the microscopic theory for transport properties has just begun. Some interesting features exhibited by the coefficient of viscosity and diffusion coefficient as a function of concentration appears to be spectacular [25]. One of the features is the sharp bending in the η vs x (or D vs x) profile around the critical concentration. Competition between the thermal excitation of ions and the variation of density with temperature is another one. In this case, for $T < T_c$, effects of atomic excitation dominate in determining the

dynamics whereas this situation reverses for $T > T_c$, that is the effect of density variation with temperatures dominates the excitation effects. The third interesting feature is the existence of a kind of scaling behavior $(\eta_c - \eta) / (T_c - T)^\beta$ with $\beta = 1.08$, near the critical temperature. These novel features showed by some segregating liquid binary alloys demand further research to understand the dynamic effects in segregating alloys.

APPENDIX

The softness parameter a_i used in the calculation are determined by fitting experimental $S(q)$ at small q (see Fig.12).

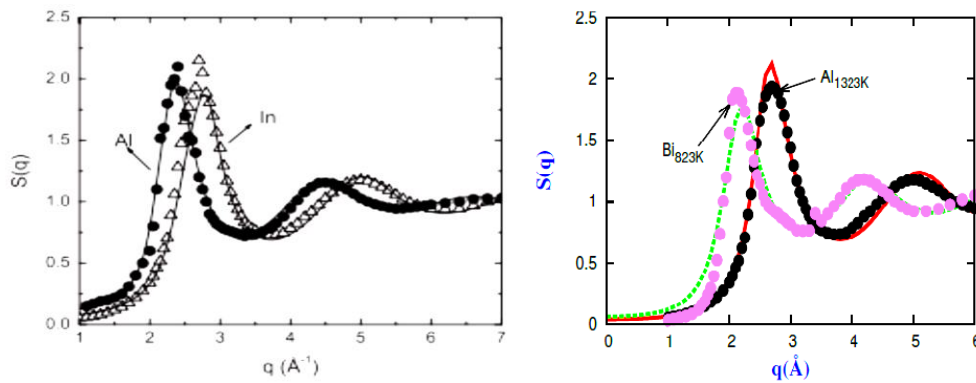


Figure 14: Determination of a_i from the best fit of $S(q)$; line theory, closed dots experiment.⁶⁴⁾ For AlIn (left) (after Faruk and Bhuiyan [21]) and for BiZn (right) (after Kasem et al. [24]).

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